

Test Bank for Chemistry The Molecular Nature of Matter 8th Jespersen

Chemistry: Molecular Nature of Matter, 8e (Jespersen)
Chapter 2 Elements, Compounds, and the Periodic Table

1) All of the following are alkali metals *except*

- A) Sr.
- B) Na.
- C) Fr.
- D) Cs.
- E) Rb.

Answer: A

Diff: 1

Section: 2.1

2) Which element is a halogen?

- A) Te
- B) O
- C) Se
- D) Uuh
- E) I

Answer: E

Diff: 1

Section: 2.1

3) Each statement accurately describes the noble gases *except* for which one?

- A) They were once known as the inert gases.
- B) He, Ne, Ar, Kr, Xe, Rn, and Uuo are part of the group.
- C) Their heavier elements do react with other elements.
- D) They belong to group VIIIA (or 18).
- E) They contain at least one metalloid.

Answer: E

Diff: 1

Section: 2.1

4) The transition metals begin in period _____ of the periodic table.

- A) 2
- B) 3
- C) 4
- D) 1
- E) 5

Answer: B

Diff: 1

Section: 2.1

5) In which family of elements does Ca belong?

- A) alkali metals
- B) alkaline earth metals
- C) halogens
- D) noble gases
- E) transition metals

Answer: B

Diff: 1

Section: 2.1

6) The elements in a specific column of the periodic table are known as

- A) a row.
- B) a period.
- C) transition elements.
- D) a group or family.
- E) transitive elements.

Answer: D

Diff: 1

Section: 2.1

7) The elements in a row of the periodic table are known as

- A) a column.
- B) a period.
- C) transition elements.
- D) a group or family.
- E) transitive elements.

Answer: B

Diff: 1

Section: 2.1

8) Which of these elements has the most similar chemical properties to sulfur?

- A) calcium
- B) oxygen
- C) phosphorus
- D) bromine
- E) nitrogen

Answer: B

Diff: 1

Section: 2.1

9) Which of these elements has the most similar chemical properties to magnesium?

- A) calcium
- B) sodium
- C) aluminum
- D) iron
- E) cesium

Answer: A

Diff: 1

Section: 2.1

10) Which of these elements has the most similar chemical properties silicon?

- A) aluminum
- B) phosphorus
- C) nitrogen
- D) silver
- E) germanium

Answer: E

Diff: 1

Section: 2.1

11) Some elements have properties that fit between metals and nonmetals. These elements are known most specifically as the

- A) metals.
- B) nonmetals.
- C) halogens.
- D) alkaline earth metals.
- E) metalloids.

Answer: E

Diff: 1

Section: 2.2

12) Which metal is a liquid at room temperature (about 25°C)?

- A) hydrogen
- B) bromine
- C) tungsten
- D) mercury
- E) chromium

Answer: D

Diff: 1

Section: 2.2

13) Which element is a gas at room temperature (about 25°C)?

- A) neon
- B) bromine
- C) tungsten
- D) mercury
- E) chromium

Answer: A

Diff: 1

Section: 2.2

14) Diamond and graphite are different forms of which element?

- A) sodium
- B) carbon
- C) mercury
- D) gold
- E) calcium

Answer: B

Diff: 1

Section: 2.2

15) Which of the following is *not* a property of metals?

- A) They have a shine called a metallic luster.
- B) They are good conductors of electricity.
- C) They are generally poor conductors of heat.
- D) They can be rolled into thin sheets.
- E) Some metals are quite hard, while some are soft.

Answer: C

Diff: 2

Section: 2.2

16) Classify the following three elements as a metal, metalloid, or nonmetal: P, Si, Al.

- A) P, metal; Si, metalloid; Al, nonmetal
- B) P, metal; Al, metalloid; Si, nonmetal
- C) Al, metal; P, metalloid; Si, nonmetal
- D) Si, metal; Al, metalloid; P, nonmetal
- E) Al, metal; Si, metalloid; P, nonmetal

Answer: E

Diff: 2

Section: 2.2

17) Classify the following three elements as a metal, metalloid, or nonmetal: Ti, S, Sb.

- A) Ti, metal; S, metalloid; Sb, nonmetal
- B) Sb, metal; S, metalloid; Ti, nonmetal
- C) S, metal; Sb, metalloid; Ti, nonmetal
- D) Sb, metal; Ti, metalloid; S, nonmetal
- E) Ti, metal; Sb, metalloid; S, nonmetal

Answer: E

Diff: 2

Section: 2.2

18) Which of these element types tend to be poor conductors of heat and electricity?

- A) metals
- B) metalloids
- C) nonmetals
- D) alkaline earth metals
- E) alkali metals

Answer: C

Diff: 1

Section: 2.2

19) Which of the following is used to represent molecular bromine?

- A) Be
- B) B
- C) 2Br
- D) Br₂
- E) Br

Answer: D

Diff: 1

Section: 2.3

20) Which combination is used to represent molecular hydrogen and atomic hydrogen, respectively?

- A) H₂, H
- B) He, H⁻
- C) H, H
- D) 2H, H⁺
- E) Hy, H

Answer: A

Diff: 2

Section: 2.3

21) Which compound exists as a diatomic molecule in its free state?

- A) magnesium
- B) manganese
- C) silicon
- D) arsenic
- E) chlorine

Answer: E

Diff: 1

Section: 2.3

22) Which compound exists as a diatomic molecule in its free state?

- A) C
- B) N
- C) Ga
- D) Ge
- E) P

Answer: B

Diff: 1

Section: 2.3

23) Which compound exists as a diatomic molecule in its free state?

- A) helium
- B) fluorine
- C) neon
- D) argon
- E) xenon

Answer: B

Diff: 1

Section: 2.3

24) Which compound is correctly represented as a hydrate?

- A) $\text{C}_2\text{H}_5\text{OH}_2^+$
- B) $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$
- C) $\text{FeH}_2(\text{CO})_4$
- D) $\text{O}_2\text{S}(\text{OH})_2$
- E) $[\text{CrCl}(\text{H}_2\text{O})_5]\text{Cl}$

Answer: B

Diff: 2

Section: 2.3

25) Which compound is correctly represented as a hydrate?

A) $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

B) $\text{HC}_2\text{H}_3\text{O}_2$

C) $\text{O}_2\text{S}(\text{OH})_2$

D) $\text{H}_2\text{S}_2\text{O}_7$

E) $\text{C}_6(\text{H}_2\text{O})_6$

Answer: A

Diff: 2

Section: 2.3

26) How many oxygen atoms are in one molecule of $\text{Mg}(\text{NO}_3)_2$?

A) 1

B) 2

C) 3

D) 5

E) 6

Answer: E

Diff: 2

Section: 2.4

27) How many oxygen atoms are in one formula unit of $\text{Mg}(\text{NO}_3)_2$?

A) 1

B) 2

C) 3

D) 5

E) 6

Answer: E

Diff: 2

Section: 2.4

28) How many hydrogen atoms are in one molecule of $\text{HC}_2\text{H}_3\text{O}_2$?

A) 1

B) 2

C) 3

D) 4

E) 5

Answer: D

Diff: 1

Section: 2.4

29) How many carbon atoms are in one molecule of C_6H_6 ?

- A) 1
- B) 2
- C) 4
- D) 6
- E) 12

Answer: D

Diff: 1

Section: 2.4

30) The number of atoms in one formula unit of the substance, $\text{CO}(\text{NH}_2)_2$, is

- A) 4
- B) 5
- C) 7
- D) 8
- E) 10

Answer: D

Diff: 2

Section: 2.4

31) The number of atoms in one formula unit of $\text{C}_2\text{H}_4(\text{COOH})_2$ is

- A) 10
- B) 11
- C) 12
- D) 14
- E) 16

Answer: D

Diff: 2

Section: 2.4

32) The number of atoms in one formula unit of the substance $\text{Cs}_2\text{SO}_4 \cdot 5\text{H}_2\text{O}$ is

- A) 4
- B) 17
- C) 22
- D) 25
- E) 33

Answer: C

Diff: 2

Section: 2.4

33) How many oxygen atoms are in one formula unit of $\text{Cu}(\text{NO}_3)_2 \cdot 5\text{H}_2\text{O}$?

- A) 2
- B) 3
- C) 5
- D) 6
- E) 11

Answer: E

Diff: 2

Section: 2.4

34) The number of atoms in one formula unit of the substance $(\text{NH}_4)_3\text{Co}(\text{CN})_6$ is

- A) 21
- B) 26
- C) 28
- D) 31
- E) 33

Answer: C

Diff: 2

Section: 2.4

35) How many atoms are there in one formula unit of $(\text{NH}_4)_4\text{Fe}(\text{CN})_6$?

- A) 15
- B) 25
- C) 28
- D) 33
- E) 35

Answer: D

Diff: 2

Section: 2.4

36) How many atoms are there in one formula unit of $\text{NiSO}_4 \cdot 7\text{H}_2\text{O}$?

- A) 9
- B) 14
- C) 27
- D) 28
- E) 33

Answer: C

Diff: 2

Section: 2.4

37) How many atoms of each element are shown in the formula $\text{H}_2\text{S}_2\text{O}_7$?

- A) 2H, 2S, 7O
- B) 1H, 2S, 4O
- C) 2H, 1S, 1O
- D) 2H, 4S, 4O
- E) 1H, 1S, 1O

Answer: A

Diff: 2

Section: 2.4

38) How many atoms of each element shown are in the formula $\text{Ni}(\text{ClO}_4)_2$?

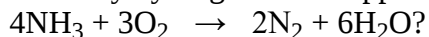
- A) 2Ni, 1Cl, 4O
- B) 1Ni, 2 Cl, 8O
- C) 2Ni, 1 Cl, 1O
- D) 2Ni, 4 Cl, 4O
- E) 1Ni, 1 Cl, 8O

Answer: B

Diff: 2

Section: 2.4

39) How many hydrogen atoms appear on the right side of the equation following equation:



- A) 2
- B) 4
- C) 6
- D) 10
- E) 12

Answer: E

Diff: 2

Section: 2.4

40) How many atoms are in one molecule of $\text{Mo}_2(\text{O}_2\text{CC}(\text{CH}_3)_3)_4$?

Hint: Remember to distribute the subscripts to all elements inside the parenthesis. Start on the inside and work your way out.

- A) 17
- B) 30
- C) 60
- D) 66
- E) 64

Answer: D

Diff: 3

Section: 2.4

41) How many atoms of each element appear on each side of the arrow in the following chemical equation? $2\text{Fe}(\text{NO}_3)_3 + 3\text{Na}_2\text{CO}_3 \rightarrow \text{Fe}_2(\text{CO}_3)_3 + 6\text{NaNO}_3$

Hint: Distribute coefficients across the entire compound for each reactant and product. Be sure to account for parenthesis.

- A) 2Fe, 6N, 18O, 6Na, 3C
- B) 2Fe, 6N, 27O, 6Na, 9C
- C) 2Fe, 6N, 27O, 6Na, 3C
- D) 2Fe, 6N, 27O, 9Na, 3C
- E) 2Fe, 6N, 21O, 18Na, 3C

Answer: C

Diff: 3

Section: 2.4

42) How many atoms of each element appear on each side of the arrow in the following chemical equation? $3\text{Cl}_3\text{BNH}_2\text{CH}_3 + 6(\text{CH}_3)_3\text{N} \rightarrow 6(\text{CH}_3)_3\text{NHCl} + \text{B}_3\text{N}_3\text{Cl}_3(\text{CH}_3)_3$

Hint: Distribute coefficients across the entire compound for each reactant and product. Be sure to account for parenthesis.

- A) 9Cl, 3B, 9N, 54H, 6C
- B) 9Cl, 3B, 9N, 69H, 21C
- C) 3Cl, 3B, 9N, 15H, 21C
- D) 9Cl, 1B, 9N, 54H, 9C
- E) 3Cl, 3B, 9N, 69H, 18C

Answer: B

Diff: 3

Section: 2.4

43) A student attempts to balance a chemical equation and comes up with the following result: $8\text{KClO}_3 + \text{C}_{12}\text{H}_{22}\text{O}_{10} \rightarrow 8\text{KCl} + 12\text{CO}_2 + 11\text{H}_2\text{O}$

However, he wrote the initial equation wrong and therefore could not balance the equation. Which element(s) are not balanced in this result?

Hint: Make a chart and identify how many elements you have on each side of the arrow.

- A) Cl
- B) O
- C) H
- D) O and H
- E) C

Answer: B

Diff: 3

Section: 2.4

44) Which of these pairs of elements would be most likely to form an ionic compound?

- A) P and Br
- B) Cu and K
- C) C and O
- D) O and Zn
- E) Al and Rb

Answer: D

Diff: 1

Section: 2.5

45) Which of these pairs of elements would be most likely to form a molecular compound?

- A) P and Br
- B) Cu and K
- C) K and O
- D) O and Zn
- E) Al and Rb

Answer: A

Diff: 1

Section: 2.5

46) The formula for the phosphate ion is

- A) PO_4^{2-} .
- B) PO_4^{3-} .
- C) PO_4^- .
- D) P_2O_4^- .
- E) $\text{P}_2\text{O}_4^{2-}$.

Answer: B

Diff: 1

Section: 2.5

47) The correct formula for the carbonate ion is

- A) $\text{C}_2\text{H}_3\text{O}_2^-$.
- B) $\text{C}_2\text{O}_4^{2-}$.
- C) CO_2^- .
- D) CO_3^- .
- E) CO_3^{2-} .

Answer: E

Diff: 1

Section: 2.5

48) When barium metal reacts with chlorine gas, it forms an ionic compound, BaCl_2 . In the course of the reaction, each Ba atom

- A) loses two protons.
- B) loses two electrons.
- C) gains two protons.
- D) gains two electrons.
- E) loses two neutrons.

Answer: B

Diff: 2

Section: 2.5

49) When barium metal reacts with chlorine gas, it forms an ionic compound, BaCl_2 . In the course of the reaction, each Cl atom

- A) loses one proton.
- B) loses one electron.
- C) gains one proton.
- D) gains one electron.
- E) loses one neutron.

Answer: D

Diff: 2

Section: 2.5

50) Write the formula for the ionic compound formed from magnesium and sulfur.

- A) MgS_2
- B) MgS
- C) Mg_2S
- D) Mg_3S_2
- E) MgS_3

Answer: B

Diff: 2

Section: 2.5

51) Write the most likely formula for the ionic compound formed from magnesium and phosphorus.

- A) MgP_2
- B) Mg_3P
- C) Mg_2P
- D) Mg_3P_2
- E) MgP_3

Answer: D

Diff: 2

Section: 2.5

52) Write the most likely formula for the ionic compound formed from calcium and selenium.

- A) CaSe
- B) Ca₂Se
- C) CaSe₂
- D) Ca₃Se
- E) CaSe₃

Answer: A

Diff: 2

Section: 2.5

53) Write the most likely formula for the ionic compound formed from magnesium and iodine.

- A) MgI
- B) Mg₂I
- C) MgI₂
- D) MgI₃
- E) Mg₃I

Answer: C

Diff: 2

Section: 2.5

54) An alkaline earth metal, which we will represent by the symbol X, reacts with a halogen, which we will represent by the symbol Q. What would be the formula of the resulting compound?

- A) XQ
- B) XQ₂
- C) XQ₄
- D) X₂Q
- E) X₄Q

Answer: B

Diff: 2

Section: 2.5

55) Aluminum reacts with a second element, which we will represent by the symbol E, to form a compound whose formula is AlE₃. Element E is most probably

- A) an actinide element.
- B) an alkali metal.
- C) a chalcogen.
- D) a halogen.
- E) a transition metal.

Answer: D

Diff: 2

Section: 2.5

56) Aluminum reacts with another element, which we will represent by the symbol Gr, to form a compound whose formula is AlGr. Element Gr is most probably

- A) an actinide element.
- B) group 2A element.
- C) a chalcogen.
- D) group 5A element.
- E) a transition metal.

Answer: D

Diff: 2

Section: 32.5

57) Bromine reacts with a metal, which we will represent by the symbol M, to form a compound whose formula is M_2Br . Element M is most probably

- A) a metalloid.
- B) a group 2A element.
- C) a chalcogen.
- D) a group 5A element.
- E) There are no elements that can react with bromine to give the formula M_2Br .

Answer: E

Diff: 2

Section: 2.5

58) Oxygen reacts with a metal, which we will represent by the symbol Wp, to form a compound whose formula is Wp_2O . Element Wp is most probably

- A) a metalloid.
- B) a group 2A element.
- C) a chalcogen.
- D) a group 1A element.
- E) there are no elements that can react with oxygen to give the formula Wp_2O .

Answer: D

Diff: 2

Section: 2.5

59) Which formula is correct because it represents a known ionic compound?

- A) Li_2Br
- B) Pb_2I_2
- C) KBr_2
- D) Rb_2Se_4
- E) Al_2S_3

Answer: E

Diff: 2

Section: 2.5

60) Which formula is correct because it represents a known ionic compound?

- A) Be_2Br_4
- B) Ca_2I_2
- C) K_2S
- D) Rb_2Br
- E) Ca_2S_3

Answer: C

Diff: 2

Section: 2.5

61) Which formula is incorrect because it does not represent a known ionic compound?

- A) BeCl_2
- B) CaI_2
- C) K_2S
- D) RbBr
- E) Ca_2O_3

Answer: E

Diff: 2

Section: 2.5

62) Which formula is incorrect because it does not represent a correctly written ionic compound?

- A) BaCl_2
- B) Al_2F_3
- C) Na_2O
- D) RbBr
- E) CaO

Answer: B

Diff: 2

Section: 2.5

63) Select the examples in which the formula for the ionic compound is not written correctly or cannot exist as written.

- I. $(\text{NH}_4)_2\text{As}$
- II. CuO
- III. $\text{Mg}(\text{C}_2\text{H}_3\text{O})_2$
- IV. $\text{Na}_3(\text{HCO}_3)_2$

- A) I only
- B) II and III
- C) III and IV
- D) I and IV
- E) I, II and IV

Answer: D

Diff: 2

Section: 2.6

64) What is the formula for the oxalate ion?

- A) CO_3^{2-}
- B) $\text{C}_4\text{O}_2^{2-}$
- C) $\text{C}_4\text{O}_4^{2-}$
- D) $\text{C}_2\text{O}_4^{2-}$
- E) $\text{C}_2\text{H}_3\text{O}_2$

Answer: D

Diff: 2

Section: 2.5

65) The formula of the compound formed from the calcium ion and acetate ion is

- A) $\text{CaC}_2\text{H}_3\text{O}_2$.
- B) $\text{Ca}_2\text{C}_2\text{H}_3\text{O}_2$.
- C) $\text{Ca}_2(\text{C}_2\text{H}_3\text{O}_2)_4$.
- D) $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$.
- E) $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_3$.

Answer: D

Diff: 2

Section: 2.5

66) Write the formula for the ionic compound formed from calcium ions and nitrate ions?

- A) Ca_3N_2
- B) $\text{Ca}(\text{NO}_3)_2$
- C) Ca_2NO_3
- D) Ca_2NO_2
- E) $\text{Ca}(\text{NO}_2)_2$

Answer: B

Diff: 2

Section: 2.5

67) The formula of the compound formed from the strontium ion and chromate ion is

- A) SrCrO_3 .
- B) SrCrO_4 .
- C) Sr_2CrO_4 .
- D) $\text{Sr}(\text{CrO}_4)_2$.
- E) $\text{Sr}_2(\text{CrO}_4)_3$.

Answer: B

Diff: 2

Section: 2.5

68) Which compound below has its formula written incorrectly?

- A) $\text{Al}(\text{H}_2\text{PO}_4)_3$
- B) $\text{Al}(\text{HCO}_3)_3$
- C) $\text{Ca}(\text{HCO}_3)_2$
- D) KHPO_4
- E) $\text{Ca}(\text{NO}_3)_2$

Answer: D

Diff: 2

Section: 2.5

69) Which compound below has its formula written incorrectly?

- A) $\text{Al}(\text{H}_2\text{CO}_3)_3$
- B) $\text{Al}(\text{H}_2\text{PO}_4)_3$
- C) $\text{Ba}(\text{HCO}_3)_2$
- D) KH_2PO_4
- E) $\text{Ca}(\text{NO}_2)_2$

Answer: A

Diff: 2

Section: 2.5

70) How many protons, neutrons, and electrons are in the ion, $^{57}\text{Fe}^{3+}$?

- A) 27 protons, 30 neutrons, and 30 electrons
- B) 26 protons, 31 neutrons, and 23 electrons
- C) 29 protons, 28 neutrons, and 26 electrons
- D) 26 protons, 31 neutrons, and 29 electrons
- E) 25 protons, 32 neutrons, and 22 electrons

Answer: B

Diff: 2

Section: 2.5

71) How many protons, neutrons, and electrons are in the ion, $^{129}\text{Te}^{2-}$?

- A) 52 protons, 77 neutrons, and 50 electrons
- B) 53 protons, 76 neutrons, and 55 electrons
- C) 52 protons, 77 neutrons, and 54 electrons
- D) 50 protons, 79 neutrons, and 52 electrons
- E) 51 protons, 78 neutrons, and 53 electrons

Answer: C

Diff: 2

Section: 2.5

72) How many protons, neutrons, and electrons are in the ion, $^{37}\text{Cl}^-$?

- A) 37 protons, 18 neutrons, and 37 electrons
- B) 18 protons, 37 neutrons, and 17 electrons
- C) 17 protons, 20 neutrons, and 18 electrons
- D) 37 protons, 20 neutrons, and 18 electrons
- E) 17 protons, 17 neutrons, and 37 electrons

Answer: C

Diff: 2

Section: 2.5

73) How many protons, neutrons, and electrons are in the cation of the compound, ^{55}MnP ?

- A) 25 protons, 30 neutrons, and 25 electrons
- B) 23 protons, 33 neutrons, and 28 electrons
- C) 28 protons, 27 neutrons, and 25 electrons
- D) 23 protons, 32 neutrons, and 26 electrons
- E) 25 protons, 30 neutrons, and 22 electrons

Answer: E

Diff: 2

Section: 2.5

74) How many protons, neutrons, and electrons are in the anion of the compound, CrCl_3^{37} ?

- A) 17 protons, 20 neutrons, and 18 electrons
- B) 18 protons, 23 neutrons, and 21 electrons
- C) 21 protons, 21 neutrons, and 16 electrons
- D) 14 protons, 23 neutrons, and 17 electrons
- E) 17 protons, 17 neutrons, and 20 electrons

Answer: A

Diff: 2

Section: 2.5

75) List the polyatomic ions, including the number of each type, present in the compound, $(\text{NH}_4)_3\text{PO}_4$.

Hint: Think about the requirements for forming ionic compounds and how this affects the number of each ion present in the formula.

- A) 4N^{3-} , 12H^+ , PO_4^{2-}
- B) 3NH_4^+ , 4PO_4^{3-}
- C) N^{3-} , 12H^+ , P^{3-} ; 4O^{2-}
- D) 3NH_4^+ , PO_4^{3-}
- E) 4NH^+ , 3PO_4^{3-}

Answer: D

Diff: 3

Section: 2.5

76) List the ions, including the number of each type, present in the compound, $\text{Fe}_2(\text{C}_2\text{O}_4)_3 \cdot 2\text{H}_2\text{O}$.

Hint: Think about the requirements for forming ionic compounds and how this affects the number of each ion present in the formula.

- A) 2Fe^{3+} , 6C^{4+} , 12O^{2-}
- B) 3Fe^{4+} , 2CO_4^{3-} ; 2OH^-
- C) 3Fe^{2+} , $3\text{C}_2\text{O}_4^{2-}$
- D) 3Fe^{4+} , 2CO_4^{3-} ; $2\text{H}_2\text{O}^{2-}$
- E) 2Fe^{3+} , 3CO_2^{2-} ; $2\text{H}_2\text{O}$

Answer: C

Diff: 3

Section: 2.5

77) Which compound below is correctly indicated as magnesium sulfate heptahydrate?

- A) $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$
- B) $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$
- C) $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
- D) $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$
- E) $7\text{MnSO}_4 \cdot 2\text{H}_2\text{O}$

Answer: B

Diff: 1

Section: 2.6

78) The name of the compound $\text{Al}(\text{SO}_4)_3$ is

- A) aluminum sulfide.
- B) aluminum sulfate.
- C) aluminum trisulfate.
- D) aluminum(III) sulfate.
- E) aluminum sulfite.

Answer: B

Diff: 2

Section: 2.6

79) What is the name of the compound, $\text{V}(\text{NO}_3)_3$?

- A) vanadium trinitrate
- B) vanadium nitrite
- C) vanadium(III) nitrite
- D) vanadium nitrate
- E) vanadium(III) nitrate

Answer: E

Diff: 2

Section: 2.6

80) What is the name for the compound, $\text{Ba}(\text{NO}_3)_2$?

- A) barium dinitrate
- B) barium dinitrite
- C) barium nitrate
- D) barium(II) nitrite
- E) barium(II) nitrate

Answer: C

Diff: 2

Section: 2.6

81) Which compound is correctly written as a hydride?

- A) $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$
- B) $\text{HC}_2\text{H}_3\text{O}_2$
- C) NaOH
- D) CaH_2
- E) $\text{C}_6\text{H}_{12}\text{O}_6$

Answer: D

Diff: 2

Section: 2.6

82) What is the name for the compound V_2O_5 ?

- A) divanadium pentoxide
- B) vanadic oxide
- C) vanadium(V) oxide
- D) vanadium(V) pentoxide
- E) vanadous oxide

Answer: C

Diff: 2

Section: 2.6

83) What is the name for the compound NaCl_3 ?

- A) sodium chlorate
- B) sodium chlorite
- C) sodium perchloride
- D) sodium trichloride
- E) There is no such compound.

Answer: E

Diff: 2

Section: 2.6

84) What is the name for the compound CuBr_2 ?

- A) copper(I) bromide(II)
- B) copper(II) bromide
- C) copper(II) bromite
- D) copper dibromide
- E) cuprous bromide

Answer: B

Diff: 2

Section: 2.6

85) What is the correct name for the compound Na_2O ?

- A) disodium oxide
- B) sodium oxide
- C) sodium(I) oxide
- D) sodium peroxide
- E) sodium superoxide

Answer: B

Diff: 2

Section: 2.6

86) Which is the correct name for the compound FeBr_3 ?

- A) ferrous bromide
- B) iron(III) bromide
- C) iron bromite
- D) iron tribromide
- E) iron tribromine

Answer: B

Diff: 2

Section: 2.6

87) What is the formula for the compound iron(II) sulfate?

- A) FeSO_4
- B) $\text{Fe}(\text{SO}_4)_2$
- C) Fe_2SO_4
- D) $\text{Fe}_2(\text{SO}_4)_3$
- E) $\text{Fe}_3(\text{SO}_4)_2$

Answer: A

Diff: 2

Section: 2.6

88) Which is a correct name for the compound Hg_2Cl_2 ?

Hint: Consider the charge on the mercury ion.

- A) dimercury dichloride
- B) mercuric chloride
- C) mercury(I) chloride
- D) mercury(II) dichloride
- E) There is no correct name; the formula should be HgCl .

Answer: C

Diff: 3

Section: 2.6

89) Which is a correct name for the compound CoF_3 ?

- A) cobalt fluoride
- B) cobalt trifluoride
- C) cobalt(III) fluoride
- D) cobaltic trifluoride
- E) cobaltous fluoride

Answer: C

Diff: 2

Section: 2.6

90) A correct name for SnF_4 is

- A) stannic tetrafluoride.
- B) stannous fluoride.
- C) stannous(IV) fluoride.
- D) tin(IV) fluoride.
- E) tin tetrafluoride.

Answer: D

Diff: 2

Section: 2.6

91) The correct formula for tin(II) nitrate is

- A) $\text{Sn}(\text{NO}_2)_2$.
- B) $\text{Sn}(\text{NO}_3)_2$.
- C) $\text{Sn}(\text{NO}_3)_3$.
- D) $\text{Sn}(\text{NO}_3)_4$.
- E) Sn_2NO_3 .

Answer: B

Diff: 2

Section: 2.6

92) What is the formula for magnesian perchlorate?

- A) MgClO_3
- B) $\text{Mg}(\text{ClO}_3)_2$
- C) Mg_2ClO_3
- D) $\text{MgO}(\text{ClO}_3)_2$
- E) There is no such compound.

Answer: E

Diff: 2

Section: 2.6

93) What is the name for $\text{Na}_2\text{Cr}_2\text{O}_7$?

- A) sodium chromium(VII)ate
- B) sodium dichromate
- C) sodium dichromium heptaoxide
- D) sodium heptaoxochromate
- E) sodium perchromate

Answer: B

Diff: 2

Section: 2.6

94) The compound $\text{Na}_2\text{S}_2\text{O}_3$ is used extensively in photographic film processing. What is its chemical name?

- A) sodium bisulfite
- B) sodium disulfur trioxide
- C) sodium oxosulfate(IV)
- D) sodium thiosulfate
- E) sodium trioxosulfite

Answer: D

Diff: 2

Section: 2.6

95) If the NtO_4^{2-} ion is called nortonate, what is the correct name for the compound K_2NtO_4 ?

- A) dipotassium nortonium tetraoxide
- B) dipotassium nortonate
- C) potassium nortonate
- D) potassium(I) nortonate
- E) potassium(II) nortonate

Answer: C

Diff: 2

Section: 2.6

96) What is the name for Cu_2SO_3 ?

- A) copper(I) sulfite
- B) copper(II) sulfite
- C) copper thiosulfate
- D) cuprous sulfate
- E) dicopper sulfur trioxide

Answer: A

Diff: 2

Section: 2.6

97) What is a correct name for the FeCrO_4 ?

- A) iron(II) chromate
- B) iron dichromate
- C) iron(III) chromium tetraoxide
- D) iron monochromate
- E) ferrous chrome

Answer: A

Diff: 2

Section: 2.6

98) What is the name for $\text{CuC}_2\text{H}_3\text{O}_2$?

- A) copper(I) acetate
- B) carbon hydrocarbonate
- C) copper monocarbonate
- D) copper(I) oxalate
- E) dicarbon acetate

Answer: A

Diff: 2

Section: 2.6

99) What is a correct name for KHCr_2O_7 ?

- A) potassium bichromite
- B) potassium bichromate
- C) potassium dichromic acid
- D) potassium monohydrogen chromite
- E) potassium monohydrogen dichromate

Answer: E

Diff: 2

Section: 2.6

- 100) What is the name for LiHPO_4 ?
- A) lithium monohydrogen phosphate
 - B) lithium hydrogen phosphoric acid
 - C) lithium hydrogen phosphorus tetraoxide
 - D) lithium monohydrogen phosphite
 - E) There is no known ionic compound with that formula.

Answer: E

Diff: 2

Section: 2.6

- 101) What is the name for Li_2HPO_4 ?
- A) lithium monohydrogen phosphate
 - B) There is no compound with that formula.
 - C) dilithium monohydrogen phosphate
 - D) lithium hydrogen phosphorus tetraoxide
 - E) lithium phosphoric acid

Answer: A

Diff: 2

Section: 2.6

- 102) What is the name for CuHSO_4 ?
- A) copper(I) hydrogen sulfate
 - B) copper(II) bisulfate acid
 - C) copper hydrogen sulfur tetraoxide
 - D) copper hydrogen sulfate
 - E) copper sulfuric acid

Answer: A

Diff: 2

Section: 2.6

- 103) Which compound is correctly indicated as cobalt(II) chloride hexahydrate?
- A) $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$
 - B) $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$
 - C) $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
 - D) $\text{NiSO}_4 \cdot 6\text{H}_2\text{O}$
 - E) $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$

Answer: A

Diff: 2

Section: 2.6

104) Which is a correct formula for mercury(I) phosphate?

Hint: Remember that mercury (I) is not a single atom.

- A) HgPO_3
- B) HgPO_4
- C) Hg_3PO_4
- D) Hg_2PO_3
- E) $(\text{Hg}_2)_3(\text{PO}_4)_2$

Answer: E

Diff: 3

Section: 2.6

105) Select the examples in which the formulas do not correctly match the names of the compounds indicated.

- | | |
|-----------------------------------|---|
| I. Sodium thiosulfate | Na_2SO_3 |
| II. Barium oxalate | BaC_2O_4 |
| III. Iron(II) sulfate hexahydrate | $\text{FeSO}_4 \cdot 6\text{H}_2\text{O}$ |
| IV. Calcium phosphate | Ca_3PO_4 |

Hint: Write the correct name of each formula without looking at the names and then compare it to the choices.

- A) II only
- B) II and III
- C) I, II and IV
- D) I and IV
- E) II, III and IV

Answer: D

Diff: 3

Section: 2.6

106) Select the examples in which the names do not correctly match the formulas of the compounds indicated.

- | | |
|--|------------------------------|
| I. $\text{Fe}_2(\text{CO}_3)_3$ | iron(III) carbonate |
| II. $\text{Cr}_2(\text{C}_2\text{O}_4)_3$ | chromium(III) oxalate |
| III. $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$ | magnesium hydrogen carbonate |
| IV. PbSO_3 | lead sulfate |

Hint: Write the correct name of each formula without looking at the names and then compare it to the choices.

- A) II only
- B) II and III
- C) III and IV
- D) I and IV
- E) I, II and IV

Answer: C

Diff: 3

Section: 2.6

107) Select the examples in which the names do not correctly match the formulas of the compounds indicated.

- | | |
|---|--------------------------------|
| I. NaClO_3 | Sodium chlorate |
| II. $(\text{NH}_4)_2\text{CO}_3$ | Ammonium(I) carbonate |
| III. $\text{Cd}(\text{H}_2\text{PO}_4)_2$ | Cadmium dihydrogen phosphate |
| IV. KMnO_4 | Potassium-manganese(VII) oxide |

Hint: Write the correct name of each formula without looking at the names and then compare it to the choices.

- A) II only
- B) II and IV
- C) III and IV
- D) I and IV
- E) I, II and IV

Answer: B

Diff: 3

Section: 2.6

108) Which compound is correctly classified as a hydrocarbon?

- A) $\text{C}_6\text{H}_{12}\text{O}_6$
- B) C_8H_{16}
- C) $\text{HC}_2\text{H}_3\text{O}_2$
- D) NaHCl
- E) $\text{C}_2\text{H}_5\text{OH}$

Answer: B

Diff: 1

Section: 2.7

109) One of the components of kerosene is an alkane with 16 carbon atoms. Which formula is an alkane?

A) $C_{16}H_{12}$

B) $C_{16}H_{22}$

C) $C_{16}H_{32}$

D) $C_{16}H_{34}$

E) $C_{16}H_{40}$

Answer: D

Diff: 2

Section: 2.7

110) The common name for the compound, CH_4 , is

A) carbon(IV) hydride.

B) carbon tetrahydride.

C) hydrocarbonate.

D) methane.

E) carbonic acid.

Answer: D

Diff: 2

Section: 2.7

111) The common name for the compound, C_2H_6 , is

A) carbon hydride.

B) carbon hexahydride.

C) ethane.

D) methane.

E) propane.

Answer: C

Diff: 2

Section: 2.7

112) Which of the following is the correct formula for the hydrocarbon hexane?

A) CH_4

B) C_2H_4

C) CH_6

D) C_6H_{14}

E) C_2H_5OH

Answer: D

Diff: 2

Section: 2.7

113) Which compound is not a simple hydrocarbon?

- A) C_6H_{12}
- B) C_8H_{16}
- C) C_2H_6
- D) $\text{C}_5\text{H}_5\text{N}$
- E) C_3H_6

Answer: D

Diff: 2

Section: 2.7

114) The formula for the compound formed between arsenic (As) and hydrogen is

- A) AsH.
- B) As_2H .
- C) AsH_2 .
- D) As_3H .
- E) AsH_3 .

Answer: E

Diff: 2

Section: 2.7

115) The most likely formula for the compound formed between antimony and chlorine is

- A) SbCl.
- B) SbCl_2 .
- C) SbCl_3 .
- D) SbCl_4 .
- E) SbCl_6 .

Answer: C

Diff: 2

Section: 2.7

116) What is the name of the compound, $\text{HI}(g)$?

- A) hydriodic acid
- B) hydrogen monoiodide
- C) hydrogen iodide
- D) iodic acid
- E) monohydrogen monoiodide

Answer: C

Diff: 1

Section: 2.8

117) What is the name of the compound, IBr_3 ?

- A) bromic iodide
- B) iodine bromate
- C) iodine tribromide
- D) iodine tribromine
- E) moniodine tribromite

Answer: C

Diff: 2

Section: 2.8

118) What is the name of the compound, S_2Cl_2 ?

- A) disulfur chlorate
- B) disulfur dichloride
- C) disulfur dichlorine
- D) sulfur(I) chloride
- E) sulfur(II) chlorine(II)

Answer: B

Diff: 2

Section: 2.8

119) What is the name of the compound, $\text{HCN}(g)$?

- A) hydrocarbonitride
- B) hydrocyanic acid
- C) hydrogen carbonitride
- D) hydrogen cyanate
- E) hydrogen cyanide

Answer: E

Diff: 2

Section: 2.8

120) A typographical error on an exam produced the formula, P_4Se_7 , in one of the questions. How would you name this compound?

- A) tetraphosphorus hexaselenide
- B) tetraphosphorus heptaselenide
- C) phosphorus heptaselenite
- D) phosphorus(IV) selenide
- E) phosphorus(VII) selenide

Answer: B

Diff: 2

Section: 2.8

121) Select the examples in which the formulas do not correctly match the names of the compounds indicated.

- | | |
|----------------------------|-------------------------|
| I. dichlorine heptoxide | Cl_2O_6 |
| II. iodine heptafluoride | I_2F_7 |
| III. dinitrogen difluoride | N_2F_2 |
| IV. tetraarsenic hexoxide | As_4O_8 |

Hint: Write the correct name of each formula without looking at the names and then compare it to the choices.

- A) II only
- B) II and III
- C) I, II and IV
- D) I and IV
- E) II, III and IV

Answer: C

Diff: 3

Section: 2.8

122) Select the examples in which the names do not correctly match the formulas of the compounds indicated.

- | | |
|-----------------------------|-------------------------|
| I. PF_5 | potassium pentafluoride |
| II. N_2O_4 | dinitrogen(IV) oxide |
| III. XeO_4 | xenon tetroxide |
| IV. Cl_2O_5 | dichlorine pentoxide |

Hint: Write the correct name of each formula without looking at the names and then compare it to the choices.

- A) II only
- B) II and III
- C) III and IV
- D) I and II
- E) I, II and IV

Answer: D

Diff: 3

Section: 2.8

123) What is the name of the compound, $\text{N}_2\text{O}_5(g)$?

- A) nitrogen oxide
- B) dinitrogen tetroxide
- C) nitrogen pentoxide
- D) dinitrogen pentoxide
- E) trinitrogen pentoxide

Answer: D

Diff: 2

Section: 2.8

124) The vertical columns in the periodic table are numbered sequentially, 1 through _____ using Arabic numerals.

Answer: 18

Diff: 1

Section: 2.1

125) Selenium is found in which group of the periodic table?

Answer: group VIA

Diff: 1

Section: 2.1

126) The group 1 elements form compounds with oxygen that dissolve in water to give solutions that

are strongly _____.

Answer: alkaline

Diff: 2

Section: 2.1

127) Which group of nonmetallic elements is called "inert"?

Answer: noble gases

Diff: 1

Section: 2.2

128) Metalloids' electrical conductivity tends to be lower than metals, but they can have higher conductivity than many metals when they are used in materials called _____.

Answer: semiconductors

Diff: 2

Section: 2.2

129) The symbol Te belongs to a metalloid in group _____.

Answer: VIA, 16

Diff: 2

Section: 2.2

130) An element is found to have a high conductivity in its pure form, is malleable, and is ductile. Based on these properties, this element would be best classified as _____.

Answer: metal

Diff: 2

Section: 2.2

131) An element is found to be a solid with low electrical conductivity but does conduct electricity. It also has a high density, is shiny, and is also brittle, shattering when hit with a hammer. Based on these properties, this element would be best classified as _____.

Answer: metalloid

Diff: 2

Section: 2.2

132) Two atoms of nitrogen combine with one atom of oxygen to form one compound, whereas two atoms of nitrogen combine with five atoms of oxygen to form another compound. The ratio of the masses of oxygen in the two compounds must be _____.

Answer: 1/5

Diff: 2

Section: 2.3

133) What formula is used to represent molecular chlorine?

Answer: Cl_2

Diff: 2

Section: 2.3

134) List the seven diatomic molecules that are the most stable form of their given element.

Answer: H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , I_2

Diff: 1

Section: 2.3

135) To show how atoms are connected in certain compounds, the chemical symbols are used to represent the atoms, and dashes are used to indicate the chemical bonds. The resulting formula is therefore referred to as _____.

Answer: a structural formula

Diff: 2

Section: 2.3

136) Write the formula for the compound that has the atoms and, or groups in the order given: 3 Fe, and two groups made up of 1 As and 4 O.

Hint: The groups of atoms are polyatomic ions.

Answer: $\text{Fe}_3(\text{AsO}_4)_2$

Diff: 3

Section: 2.3

137) Write the formula for the hydrated compound that has the atoms and, or groups in the order given: 1 K, 1 Al, two groups of 1 S and 4 O, and twelve groups made up of 2 H and 1 O.

Hint: The groups of atoms are clusters such as polyatomic ions or water.

Answer: $\text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}$

Diff: 3

Section: 2.3

138) List how many of each type of element are present in one molecule of sucrose, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$.

Answer: 12 carbon, 22 hydrogen, 11 oxygen

Diff: 1

Section: 2.4

139) List how many oxygen atoms are present in one molecule of H_3PO_4

Answer: 4

Diff: 1

Section: 2.4

140) List how many oxygen atoms are present in one formula unit of $\text{CoSO}_4 \cdot 6\text{H}_2\text{O}$.

Answer: 10

Diff: 2

Section: 2.4

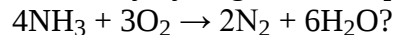
141) How many hydrogen atoms are present in the formula, $(\text{NH}_4)_3\text{PO}_4$?

Answer: 12

Diff: 2

Section: 2.4

142) How many hydrogen atoms appear on the reactant side of the equation,



Answer: 12

Diff: 2

Section: 2.4

143) How many of each type of atoms are needed on the left to balance the equation?



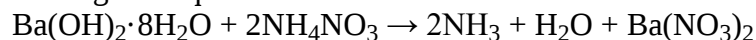
Hint: Distribute coefficients across the entire compound and remember how subscripts apply to everything inside the parentheses.

Answer: 2Al, 6H, 6O

Diff: 3

Section: 2.4

144) How many additional hydrogen atoms and oxygen atoms are required on the right side to balance the given equation?



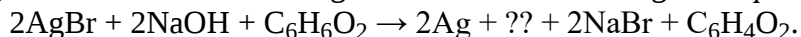
Hint: Distribute coefficients across the entire compound and remember how subscripts apply to everything inside the parentheses.

Answer: 18H, 9O

Diff: 3

Section: 2.4

145) What molecule is missing that would balance the given equation?



Balance the equation by entering the correct coefficient and formula for the missing molecule.

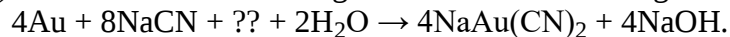
Hint: Figure out which elements are missing and consider the ratio of those elements.

Answer: $2\text{H}_2\text{O}$

Diff: 3

Section: 2.4

146) What molecule is missing that would balance the given equation?



If a coefficient other than one is needed provide the coefficient.

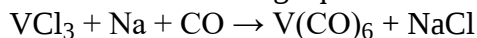
Hint: Figure out which elements are missing and consider the ratio of those elements.

Answer: O_2

Diff: 3

Section: 2.4

147) Balance the following equation:

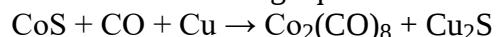


Answer: $\text{VCl}_3 + 3\text{Na} + 6\text{CO} \rightarrow \text{V(CO)}_6 + 3\text{NaCl}$

Diff: 2

Section: 2.4

148) Balance the following equation:



Answer: $2\text{CoS} + 8\text{CO} + 4\text{Cu} \rightarrow \text{Co}_2(\text{CO})_8 + 2\text{Cu}_2\text{S}$

Diff: 2

Section: 2.4

149) What is the charge on all the ions of metals of Group IIA?

Answer: 2^+

Diff: 1

Section: 2.5

150) What is the charge on all the ions of non-metals of Group VIIA?

Answer: -1

Diff: 1

Section: 2.5

151) What is the formula for the sulfide ion?

Answer: S^{2-}

Diff: 1

Section: 2.5

152) How many protons and electrons are in the N^{3-} ion?

Answer: 7 protons; 10 electrons

Diff: 2

Section: 2.5

153) How many protons and electrons are in the S^{2-} ion?

Answer: 16 protons; 18 electrons

Diff: 2

Section: 2.5

154) How many protons and electrons are in the Ca^{2+} ion?

Answer: 20 protons; 18 electrons

Diff: 2

Section: 2.5

155) How many electrons are lost when aluminum forms a cation?

Answer: 3

Diff: 2

Section: 2.5

156) How many electrons are lost when zinc forms a cation?

Answer: 2

Diff: 2

Section: 2.5

157) How many electrons are gained when sulfur forms an anion?

Answer: 2

Diff: 2

Section: 2.5

158) The correct formula for the compound formed from the lithium ion and PO_4^{3-} is _____.

Answer: Li_3PO_4

Diff: 2

Section: 2.5

159) The formula for the compound formed from the barium ion and SO_3^{2-} is _____.

Answer: BaSO_3

Diff: 2

Section: 2.5

160) The formula formed from the calcium ion and ClO_2^- is _____.

Answer: $\text{Ca}(\text{ClO}_2)_2$

Diff: 2

Section: 2.5

161) What is the formula of the compound formed from Cr^{3+} and H_2PO_4^- ?

Answer: $\text{Cr}(\text{H}_2\text{PO}_4)_3$

Diff: 2

Section: 2.5

162) What is the formula of the compound formed from the calcium ion and HCO_3^- ?

Answer: $\text{Ca}(\text{HCO}_3)_2$

Diff: 2

Section: 2.5

163) What is the name of the following compound: $(\text{NH}_4)_2\text{SO}_4$?

Answer: ammonium sulfate

Diff: 2

Section: 2.5

164) What is the name of the following compound: $\text{Cr}_2(\text{SO}_4)_3$?

Answer: chromium(III) sulfate

Diff: 2

Section: 2.6

165) What is the name of the following compound: $\text{V}_3(\text{PO}_4)_4$?

Answer: vanadium(IV) phosphate

Diff: 2

Section: 2.6

166) What is the name of the following compound: Mn_2O_7 ?

Answer: manganese(VII) oxide

Diff: 2

Section: 2.6

167) What is the name of the following compound: NH_4NO_3 ?

Answer: ammonium nitrate

Diff: 2

Section: 2.6

168) What is the name of the following compound: $\text{Ba}(\text{OH})_2$?

Answer: barium hydroxide

Diff: 2

Section: 2.6

169) What is the name of the following compound: KHCO_3 ?

Answer: potassium hydrogen carbonate or potassium bicarbonate

Diff: 2

Section: 2.6

170) What is the correct formula for lead(IV) chloride?

Answer: PbCl_4

Diff: 2

Section: 2.6

171) What is the correct formula for calcium phosphate?

Answer: $\text{Ca}_3(\text{PO}_4)_2$

Diff: 2

Section: 2.6

172) What is the correct formula for magnesium sulfate?

Answer: MgSO_4

Diff: 2

Section: 2.6

173) What is the correct formula for sodium sulfide?

Answer: Na_2S

Diff: 2

Section: 2.6

174) What is the correct formula for chromium(VI) oxide?

Answer: CrO_3

Diff: 2

Section: 2.6

175) Predict the formula of the compound formed between chlorine and hydrogen.

Answer: HCl

Diff: 1

Section: 2.7

176) Predict the formula of the compound formed between sulfur and hydrogen.

Answer: H_2S

Diff: 2

Section: 2.7

177) Hydrocarbons are organic compounds which have general formula _____.

Answer: $\text{C}_n\text{H}_{2n+2}$

Diff: 2

Section: 2.7

178) Write the formula of the alkane hydrocarbon with seven carbon atoms.

Answer: C_7H_{16}

Diff: 2

Section: 2.7

179) What is the name for the compound PBr_3 ?

Answer: phosphorus tribromide

Diff: 1

Section: 2.8

180) What is the name for the compound Si_3N_4 ?

Answer: trisilicon tetranitride

Diff: 1

Section: 2.8

181) The name for As_4S_{10} is _____.

Answer: tetraarsenic decasulfide

Diff: 2

Section: 2.8

182) What is the formula for dichlorine heptoxide?

Answer: Cl_2O_7

Diff: 2

Section: 2.8

183) The alkali metals like sodium and potassium are soft metals, so they are unreactive towards water.

Answer: FALSE

Diff: 2

Section: 2.1

184) The number of protons in the nucleus of an atom, determines the order of elements in the periodic table.

Answer: TRUE

Diff: 2

Section: 2.1

185) Due to their properties, and where they exist in nature, the group 2A metals are called the lanthanides.

Answer: FALSE

Diff: 2

Section: 2.1

186) Elements that are part of the actinides are all composed of radioactive gases.

Answer: FALSE

Diff: 2

Section: 2.1

187) Some of the nonmetals are solids at room temperature.

Answer: TRUE

Diff: 1

Section: 2.2

188) Metalloids are capable of conducting an electric current.

Answer: TRUE

Diff: 1

Section: 2.2

189) From left to right, across a row on the periodic table, there is a gradual change in properties from nonmetallic to metallic properties.

Answer: FALSE

Diff: 1

Section: 2.2

190) Metalloids tend to be malleable and ductile in nature.

Answer: FALSE

Diff: 1

Section: 2.2

191) The formula, N_2 , is used to represent elemental nitrogen.

Answer: TRUE

Diff: 1

Section: 2.3

192) When interpreting the formula, $\text{CO}(\text{NH}_2)_2$, it should be noted that the group of atoms within the parentheses occurs twice.

Answer: TRUE

Diff: 1

Section: 2.3

193) An important characteristic of a compound's formula is it specifies the atomic composition of the compound.

Answer: TRUE

Diff: 2

Section: 2.3

194) When iron and sulfur combine chemically, the properties of the resulting compound are similar to that of each of the elements.

Answer: FALSE

Diff: 1

Section: 2.4

195) Four molecules of the only product formed in the incomplete equation below are needed to ensure that the equation is balanced. $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow ??$

Answer: TRUE

Diff: 2

Section: 2.4

196) Ionic compounds are generally formed when metals react with nonmetals.

Answer: TRUE

Diff: 1

Section: 2.5

197) The phosphide ion has 18 electrons and 18 protons.

Answer: FALSE

Diff: 2

Section: 2.5

198) The subscripts in the formulas do not normally produce an electrically neutral formula unit in ionic compounds.

Answer: FALSE

Diff: 2

Section: 2.5

199) The name of MnCl_3 is magnesium(III) chloride.

Answer: FALSE

Diff: 1

Section: 2.6

200) It is important to specify how many cations and anions are present in ionic compounds.

Answer: FALSE

Diff: 2

Section: 2.6

201) The formula for magnesium phosphide is Mg_3P_2 .

Answer: TRUE

Diff: 2

Section: 2.6

202) As a general rule, molecular compounds are formed when nonmetallic elements combine.

Answer: TRUE

Diff: 1

Section: 2.7

203) The elements, carbon and oxygen, can combine to form only one compound.

Answer: FALSE

Diff: 1

Section: 2.7

204) Phosphorus can combine with hydrogen to form the compound, PH_3 .

Answer: TRUE

Diff: 1

Section: 2.7

205) The compound N_2O_4 is named nitrate tetraoxide.

Answer: FALSE

Diff: 1

Section: 2.8

206) The name for ZnBr_2 , is zirconium bromide.

Answer: FALSE

Diff: 1

Section: 2.8

207) A name for the compound P_4Se_{10} is phosphorus(IV) selenium.

Answer: FALSE

Diff: 2

Section: 2.8

208) A name for CrBr_2 , is chromic bromide.

Answer: FALSE

Diff: 2

Section: 2.8

209) The name for RbClO_4 , is rubidium(I) perchlorate.

Answer: FALSE

Diff: 2

Section: 2.8

210) A name for $\text{Ni}(\text{OCl})_2$ is nickel(II) hypochlorite.

Answer: TRUE

Diff: 2

Section: 2.8

211) A name for the compound $\text{Mn}(\text{ClO}_4)_2$, is magnesium chlorate.

Answer: FALSE

Diff: 2

Section: 2.8

212) The name for $\text{K}_2\text{Cr}_2\text{O}_7$, is potassium dichromium heptaoxide.

Answer: FALSE

Diff: 2

Section: 2.8

213) A compound is known to contain one C atom for each water molecule (H_2O). If the compound has six carbon atoms, what is the general formula representing the compound?

Answer: $\text{C}_6\text{H}_{12}\text{O}_6$

Diff: 2

Section: 2.3

214) How many silicon and oxygen atoms are in the formula, $\text{Ca}_3\text{Mg}_5(\text{Si}_4\text{O}_{11})_2(\text{OH})_2$?

A) 3 Si, 5 O

B) 8 Si, 24 O

C) 4 Si, 11 O

D) 2 Si, 2 O

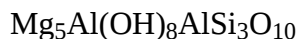
E) 5 Si, 3 O

Answer: B

Diff: 2

Section: 2.3

215) What is the total number of atoms represented by the following formula?



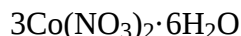
- A) 36
- B) 28
- C) 8
- D) 24
- E) 42

Answer: A

Diff: 2

Section: 2.3

216) What is the total number of atoms represented by the following?



- A) 35
- B) 28
- C) 8
- D) 81
- E) 42

Answer: D

Diff: 2

Section: 2.3

217) Through analysis it was found that an unknown molecule contains 19.8 g of nitrogen for every 65.0 g of the molecule. How many grams of nitrogen would 1.35 grams of the molecule contain?

Hint: Use the masses of the elements to find the ratio of the elements.

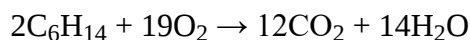
- A) 0.305 g
- B) 3.28 g
- C) 0.411 g
- D) 0.0681 g
- E) 0.226 g

Answer: C

Diff: 3

Section: 2.3

218) What is the total number of atoms reacting in the chemical reaction below?



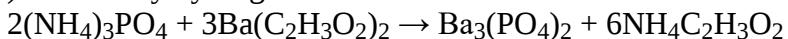
- A) 35
- B) 82
- C) 41
- D) 78
- E) 21

Answer: D

Diff: 2

Section: 2.4

219) How many hydrogen atoms are on the reactant side of the chemical equation below?



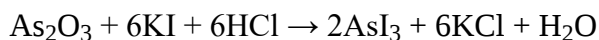
- A) 35
- B) 28
- C) 8
- D) 24
- E) 42

Answer: E

Diff: 2

Section: 2.4

220) What single coefficient is needed on the water formula to balance the following chemical equation?



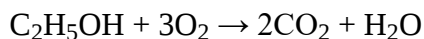
- A) 2
- B) 3
- C) 4
- D) 5
- E) 6

Answer: B

Diff: 2

Section: 2.4

221) What single coefficient is needed on the water formula to balance the following chemical equation?



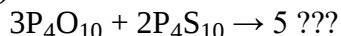
- A) 2
- B) 3
- C) 4
- D) 5
- E) 6

Answer: B

Diff: 2

Section: 2.4

222) Write the formula of the single product in the reaction below if its coefficient is 5.



Hint: Determine the necessary ratio of elements on the reactant side.

- A) $\text{P}_6\text{O}_6\text{S}_5$
- B) $\text{P}_4\text{O}_6\text{S}_4$
- C) $\text{P}_4\text{O}_{16}\text{S}_6$
- D) $\text{P}_8\text{O}_3\text{S}_8$
- E) $\text{P}_{12}\text{O}_3\text{S}_{10}$

Answer: B

Diff: 3

Section: 2.4

223) Two elements, Qr and E, combine to form an ionic compound whose formula is QrE_2 . Qr also combines with element Z to form an ionic compound, Qr_3Z_2 . Based on this information, what is a reasonable value for the charge on E? (Assume that Qr has the same charge in both compounds.)

Hint: Consider the rules for forming ionic compounds to figure out what the possible charge is. Remember ionic compounds are neutral.

- A) 1^+
- B) 1
- C) 2^+
- D) 2
- E) 3

Answer: B

Diff: 3

Section: 2.5

224) Two elements, Qr and Z, combine to form an ionic compound containing simple ions whose formula is Qr_2Z_3 . Calcium also combines with element Z to form an ionic compound containing simple ions whose formula is CaZ . Qr combines with a third element, E, to form an ionic compound containing simple ions whose formula is QrE_3 . Based on this information, what is a reasonable formula for the compound formed when magnesium combines with element E to form a simple ionic compound? (Assume that Qr has the same charge in both compounds.)

Hint: Consider the rules for forming ionic compounds to figure out what the possible charge is. Remember ionic compounds are neutral.

- A) MgE
- B) Mg_2E
- C) MgE_2
- D) Mg_2E_3
- E) Mg_3E_2

Answer: C

Diff: 3

Section: 2.5

225) Vitellium phosphate has the formula, $\text{Vi}_3(\text{PO}_4)_2$, while sodium nortonate has the formula, Na_2NtO_4 . Which of the following would be the expected formula for vitellium nortonate? (Imaginary elements are used in this question.)

Hint: Consider the rules for forming ionic compounds to figure out what the possible charge is. Remember ionic compounds are neutral.

- A) ViNtO_4
- B) Vi_2NtO_4
- C) $\text{Vi}(\text{NtO}_4)_2$
- D) $\text{Vi}_2(\text{NtO}_4)_3$
- E) $\text{Vi}_3(\text{NtO}_4)_2$

Answer: A

Diff: 3

Section: 2.5

226) Engrium sulfate has the formula, $\text{En}_2(\text{SO}_4)_3$, while sodium nortonite has the formula Na_2NtO_3 . Based on these names and formulas, what would you expect for the formula of engrium nortonate? (Imaginary elements are used in this question.)

Hint: Consider the rules for forming ionic compounds to figure out what the possible charge is. Remember ionic compounds are neutral.

- A) EnNtO_4
- B) En_2NtO_4
- C) $\text{En}(\text{NtO}_4)_2$
- D) $\text{En}_2(\text{NtO}_4)_3$
- E) $\text{En}_3(\text{NtO}_4)_2$

Answer: D

Diff: 3

Section: 2.5

227) What is the name for the ionic compound CuH_2CrO_4 ?

- A) copper(I) hydrogen chromate
- B) copper(II) bichromic acid
- C) copper hydrogen chromate tetraoxide
- D) copper hydrogen sulfate
- E) There is no known ionic compound with this formula.

Answer: E

Diff: 2

Section: 2.6

228) What is the formula for manganese(III) monohydrogen phosphate?

- A) MnHO_4
- B) MnHPO_4
- C) MnHPO_3
- D) $\text{Mn}_2(\text{HPO}_4)_3$
- E) Mn_3HPO_4

Answer: D

Diff: 2

Section: 2.6

229) What is the formula for cobalt(III) dihydrogen phosphate?

- A) Co_3HPO_4
- B) $\text{Co}_2\text{H}(\text{PO}_4)_3$
- C) $\text{Co}(\text{H}_2\text{PO}_4)$
- D) $\text{Co}(\text{H}_2\text{PO}_4)_3$
- E) CoH_3PO_4

Answer: D

Diff: 2

Section: 2.6

230) What are the likely formulas of three different hydrocarbons each with 12 hydrogen atoms?

- A) $\text{C}_{12}\text{H}_{12}$; C_6H_{12} ; C_5H_{12}
- B) $\text{C}_{12}\text{H}_{12}$; C_7H_{12} ; CH_{12}
- C) C_5H_{12} ; C_6H_{12} ; C_7H_{12}
- D) $\text{C}_{12}\text{H}_{12}$; C_9H_{12} ; CH_{12}
- E) CH_{12} ; $\text{C}_{12}\text{H}_{12}$; C_8H_{12}

Answer: C

Diff: 2

Section: 2.7

231) What are the likely formulas of three different hydrocarbons each with 5 carbon atoms?

- A) C_5H_2 ; C_5H_{12} ; C_5H_{14}
- B) C_5H_{12} ; C_5H_{10} ; C_5H_8
- C) C_5H_{12} ; C_5H_{19} ; C_5H_{11}
- D) C_5H_{12} ; C_5H_{19} ; C_5H_{30}
- E) C_5H_3 ; C_5H_{12} ; C_5H_{22}

Answer: B

Diff: 2

Section: 2.7

232) Starting with the hydrocarbon, C_6H_{14} , what is the most likely formula of the alcohol formed from this hydrocarbon?

- A) $C_6H_{15}(OH)_2$
- B) C_6HOH
- C) $C_5H_{14}OH$
- D) $C_6H_{13}OH$
- E) $C_5H_{12}OH$

Answer: D

Diff: 2

Section: 2.7

233) What is the most likely name for BrF ?

- A) bromine monofluoride
- B) bromine fluorine
- C) monobromide fluoride
- D) bromine difluorine
- E) bromide fluorine

Answer: A

Diff: 2

Section: 2.8

234) What is the best name for the I_2O_5 molecule?

- A) diiodine pentoxide
- B) iodine pentoxygen
- C) pentoxygen iodide
- D) iodine dioxide
- E) diiodide oxide

Answer: A

Diff: 2

Section: 2.8

235) What is the most likely name for IF_7 ?

- A) diiodine pentafluoride
- B) iodine heptafluoride
- C) pentafluorine iodide
- D) iodine fluoride
- E) diiodide hexafluoride

Answer: B

Diff: 2

Section: 2.8

236) What is the formula for the compound named hydrogen sulfide?

Answer: H_2S

Diff: 2

Section: 2.8

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