

Test Bank for Introductory Chemistry - An Atoms First Approach 2nd Burdge

Introductory Chemistry, 2e (Burdge)
Chapter 2 Electrons and the Periodic Table

1) Visible light, radio waves, microwave radiation, infrared, ultraviolet radiation, X-rays, and gamma rays all constitute the electromagnetic spectrum. Which of the following characteristics do all of these kinds of radiation share?

- A) They all have the ability to generate heat in objects.
- B) They all have the same frequencies.
- C) They are all the transmission of energy in the form of waves.
- D) They have equal energies.
- E) They have the same electron spin state.

Answer: C

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure

Bloom's: 3. Apply

Subtopic: Electromagnetic Radiation (Wave Properties)

Accessibility: Keyboard Navigation

Gradable: automatic

2) Select the arrangement of electromagnetic radiation which starts with the shortest wavelength and increases to longest wavelength.

- A) radio, infrared, ultraviolet, gamma rays
- B) radio, ultraviolet, infrared, gamma rays
- C) gamma rays, radio, ultraviolet, infrared
- D) gamma rays, infrared, radio, ultraviolet
- E) gamma rays, ultraviolet, infrared, radio

Answer: E

Difficulty: 3 Hard

Topic: Quantum Theory and Atomic Structure

Bloom's: 5. Evaluate

Subtopic: Electromagnetic Radiation (Wave Properties); Scientific Notation and Significant Figures; Measurement (SI Units)

Accessibility: Keyboard Navigation

Gradable: automatic

3) Select the arrangement of electromagnetic radiation which starts with the lowest energy and increases to the greatest energy.

- A) radio, infrared, ultraviolet, gamma rays
- B) radio, ultraviolet, infrared, gamma rays
- C) gamma rays, infrared, radio, ultraviolet
- D) gamma rays, ultraviolet, infrared, radio
- E) infrared, ultraviolet, radio, gamma rays

Answer: A

Difficulty: 3 Hard

Topic: Quantum Theory and Atomic Structure

Bloom's: 5. Evaluate

Subtopic: Electromagnetic Radiation (Wave Properties); Scientific Notation and Significant Figures; Measurement (SI Units)

Accessibility: Keyboard Navigation

Gradable: automatic

4) What is the emission of light at only specific wavelengths?

- A) emission spectra
- B) hydrogen spectrum
- C) wave spectra
- D) limited spectra
- E) line spectra

Answer: E

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Electromagnetic Radiation (Wave Properties); Scientific Notation and Significant Figures; Measurement (SI Units); Atomic Spectra (Bohr Model of the Atom)

Accessibility: Keyboard Navigation

Gradable: automatic

5) List the following types of radiation from lowest frequency to highest frequency: microwave, X ray, ultraviolet, visible, and infrared

- A) microwave < infrared < visible < ultraviolet < X ray
- B) X ray < ultraviolet < visible < infrared < microwave
- C) visible < ultraviolet < microwave < X ray < infrared
- D) infrared < X ray < microwave < ultraviolet < visible
- E) infrared < visible < microwave < ultraviolet < X ray

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 5. Evaluate

Subtopic: Electromagnetic Radiation (Wave Properties)

Accessibility: Keyboard Navigation

Gradable: automatic

6) Which of the following electron transitions would be expected to emit any light in the Bohr model of the atom?

- A) $n = 1$ to $n = 3$
- B) $n = 5$ to $n = 6$
- C) $n = 2$ to $n = 5$
- D) $n = 4$ to $n = 3$

Answer: D

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Atomic Spectra (Bohr Model of the Atom)

Accessibility: Keyboard Navigation

Gradable: automatic

7) Which of the following electron transitions would be expected to emit any light in the Bohr model of the atom?

- A) $n = 1$ to $n = 4$
- B) $n = 3$ to $n = 1$
- C) $n = 2$ to $n = 3$
- D) $n = 5$ to $n = 7$

Answer: B

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Atomic Spectra (Bohr Model of the Atom)

Accessibility: Keyboard Navigation

Gradable: automatic

8) Which of the following electron transitions would be expected to absorb any light in the Bohr model of the atom?

- A) $n = 1$ to $n = 3$
- B) $n = 3$ to $n = 2$
- C) $n = 4$ to $n = 2$
- D) $n = 6$ to $n = 5$

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Atomic Spectra (Bohr Model of the Atom)

Accessibility: Keyboard Navigation

Gradable: automatic

9) Which of the following electron transitions would be expected to absorb any light in the Bohr model of the atom?

- A) $n = 7$ to $n = 2$
- B) $n = 5$ to $n = 6$
- C) $n = 1$ to $n = 3$
- D) $n = 3$ to $n = 5$

Answer: C

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Atomic Spectra (Bohr Model of the Atom)

Accessibility: Keyboard Navigation

Gradable: automatic

10) The size of an atomic orbital is associated with

- A) the principal quantum number (n).
- B) the angular momentum quantum number (l).
- C) the magnetic quantum number (m_l).
- D) the spin quantum number (m_s).
- E) the angular momentum and magnetic quantum numbers, together.

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Components of Matter

Bloom's: 2. Understand

Subtopic: Quantum Numbers; Atomic Theories

Accessibility: Keyboard Navigation

Gradable: automatic

- 11) Atomic orbitals developed using quantum mechanics
- A) describe regions of space in which one is most likely to find an electron.
 - B) describe exact paths for electron motion.
 - C) give a description of the atomic structure, which is essentially the same as the Bohr model.
 - D) allow scientists to calculate an exact volume for the hydrogen atom.
 - E) are in conflict with the Heisenberg uncertainty principle.

Answer: A

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Components of Matter

Bloom's: 3. Apply

Subtopic: Quantum Numbers; Atomic Theories

Accessibility: Keyboard Navigation

Gradable: automatic

- 12) The number of orbitals in a *d* subshell is
- A) 1.
 - B) 2.
 - C) 3.
 - D) 5.
 - E) 7.

Answer: D

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

- 13) How many orbitals can have the $3p$ description in a given atom?
- A) 1
 - B) 2
 - C) 3
 - D) 5

Answer: C

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

14) How many orbitals can have the $3d$ description in a given atom?

- A) 1
- B) 2
- C) 3
- D) 5

Answer: D

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

15) How many orbitals can have the $4s$ description in a given atom?

- A) 1
- B) 2
- C) 3
- D) 5

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

16) How many orbitals can have the $4p$ description in a given atom?

- A) 1
- B) 2
- C) 3
- D) 4

Answer: C

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

17) Determine which sublevel designation is legitimate.

- A) $1f$
- B) $2d$
- C) $3c$
- D) $4s$

Answer: D

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

18) Determine which sublevel designation is legitimate.

- A) $1p$
- B) $2p$
- C) $3f$
- D) $4z$

Answer: B

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

19) Determine which sublevel designation is not legitimate.

- A) $1p$
- B) $2s$
- C) $3d$
- D) $4p$

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

20) Determine which sublevel designation is not legitimate.

- A) 4s
- B) 2d
- C) 3s
- D) 5p

Answer: B

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

21) How many orbitals are there in the $n = 4$ level of the H-atom?

- A) 4
- B) 6
- C) 8
- D) 16
- E) 18

Answer: D

Difficulty: 3 Hard

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 3. Apply

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

22) Each shell (principal energy level) of quantum number n contains n subshells.

Answer: TRUE

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure

Bloom's: 2. Understand

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

23) For all atoms of the same element, the 2s orbital is larger than the 1s orbital.

Answer: TRUE

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure

Bloom's: 3. Apply

Subtopic: Quantum Numbers

Accessibility: Keyboard Navigation

Gradable: automatic

24) The orbital diagram for a ground-state nitrogen atom is

	1s	2s		2p	
A	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	$\uparrow$	$\uparrow$
B	$\uparrow\downarrow$	$\uparrow$	$\uparrow\downarrow$	$\uparrow$	—
C	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	—
D	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	$\uparrow$
E	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$

A) A

B) B

C) C

D) D

E) E

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Gradable: automatic

25) The orbital diagram for a ground-state oxygen atom is

	1s	2s	2p		
A	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	$\uparrow$	$\uparrow$
B	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	—
C	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	—
D	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	$\uparrow$
E	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$

- A) A
- B) B
- C) C
- D) D
- E) E

Answer: D

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Gradable: automatic

26) The orbital diagram for a ground-state carbon atom is

	1s	2s	2p		
A	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	—	—
B	$\uparrow\downarrow$	$\uparrow$	$\uparrow$	$\uparrow$	$\uparrow$
C	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	$\uparrow$	$\uparrow$
D	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$	$\uparrow$	—
E	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow$

- A) A
- B) B
- C) C
- D) D
- E) E

Answer: D

Difficulty: 1 Easy

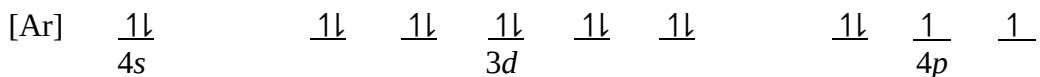
Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Gradable: automatic

27) Which ground-state atom has an electron configuration described by the following *orbital diagram*?



- A) phosphorus
- B) germanium
- C) selenium
- D) tellurium
- E) potassium

Answer: C

Difficulty: 1 Easy

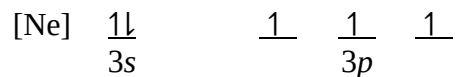
Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Gradable: automatic

28) Which ground-state atom has an electron configuration described by the following *orbital diagram*?



- A) phosphorus
- B) nitrogen
- C) arsenic
- D) vanadium
- E) sulfur

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Gradable: automatic

29) How many unpaired electrons does a ground-state atom of sulfur have?

- A) 0
- B) 1
- C) 2
- D) 3
- E) 4

Answer: C

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Pauli Exclusion Principle

Accessibility: Keyboard Navigation

Gradable: automatic

30) Which element has the following ground-state electron configuration?

$1s^2 2s^2 2p^6 3s^2$

- A) Na
- B) Mg
- C) Al
- D) Si
- E) Ne

Answer: B

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

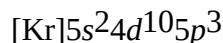
Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Hund's Rule; Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

31) Which element has the following ground-state electron configuration?



- A) Sn
- B) Sb
- C) Pb
- D) Bi
- E) Te

Answer: B

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

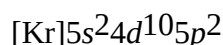
Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle);
Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

32) Which element has the following ground-state electron configuration?



- A) Sn
- B) Sb
- C) Pb
- D) Ge
- E) Te

Answer: A

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle);
Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

33) The electron configuration of a ground-state Co atom is

- A) $[\text{Ar}]4s^23d^7$.
- B) $1s^22s^22p^63s^23d^9$.
- C) $[\text{Ne}]3s^23d^7$.
- D) $[\text{Ar}]4s^13d^5$.
- E) $[\text{Ar}]4s^24d^7$.

Answer: A

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

34) The electron configuration of a ground-state vanadium atom is

- A) $[\text{Ar}]4s^24d^3$.
- B) $[\text{Ar}]4s^24p^3$.
- C) $[\text{Ar}]4s^23d^3$.
- D) $[\text{Ar}]3d^5$.
- E) $[\text{Ar}]4s^23d^7$.

Answer: C

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

35) The ground-state electron configuration for an atom of indium is

- A) $[\text{Kr}]5s^24p^64d^5$.
- B) $[\text{Ar}]4s^23d^{10}4p^1$.
- C) $[\text{Ar}]4s^24p^63d^5$.
- D) $[\text{Kr}]5s^25p^64d^5$.
- E) $[\text{Kr}]5s^24d^{10}5p^1$.

Answer: E

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

36) The ground-state electron configuration of a calcium atom is

- A) $[\text{Ne}]3s^2$.
- B) $[\text{Ne}]3s^23p^6$.
- C) $[\text{Ar}]4s^13d^1$.
- D) $[\text{Ar}]4s^2$.
- E) $[\text{Ar}]3d^2$.

Answer: D

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

37) Select the correct electron configuration for sulfur ($Z = 16$).

- A) $1s^2 1p^6 2s^2 2p^6$
- B) $1s^2 2s^2 2p^8 3s^2 3p^4$
- C) $1s^2 2s^2 2p^8 3s^2 3p^2$
- D) $1s^2 2s^2 2p^6 3s^2 3p^4$
- E) $1s^2 2s^2 2p^6 3s^2 3d^4$

Answer: D

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

38) Select the correct electron configuration for Te ($Z = 52$).

- A) $[\text{Kr}]5s^2 5p^6 4d^8$
- B) $[\text{Kr}]5s^2 5d^{10} 5p^4$
- C) $[\text{Kr}]5s^2 4d^{10} 5p^6$
- D) $[\text{Kr}]5s^2 4f^{14}$
- E) $[\text{Kr}]5s^2 4d^{10} 5p^4$

Answer: E

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle);

Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

39) What is the correct electron configuration for a germanium (Ge) atom?

- A) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4p^2$
- B) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^2$
- C) $1s^2 2s^2 2p^6 3s^2 3p^2$
- D) $1s^2 2s^2 3s^2 3p^5$
- E) None of the answers is correct.

Answer: B

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle);
Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

40) The electronic structure $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^8$ refers to the ground state of

- A) Kr.
- B) Ni.
- C) Fe.
- D) Pd.
- E) None of the answers is correct.

Answer: B

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle);
Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

41) How many electrons are in the $4p$ orbitals of selenium?

- A) 0
- B) 2
- C) 4
- D) 5
- E) 6

Answer: C

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

42) How many electrons are in the $4p$ orbitals of vanadium?

- A) 0
- B) 2
- C) 4
- D) 5
- E) 6

Answer: A

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

43) How many electrons are in the $4d$ orbitals of Tc?

- A) 1
- B) 2
- C) 3
- D) 4
- E) 5

Answer: E

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

44) How many electrons are there in the 2nd principal energy level ($n = 2$) of a phosphorus atom?

- A) 3
- B) 5
- C) 6
- D) 8
- E) 10

Answer: D

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

45) How many electrons are there in the 3rd principal energy level ($n = 3$) of a phosphorus atom?

- A) 3
- B) 5
- C) 6
- D) 8
- E) 10

Answer: B

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

46) What element is represented by the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^5$?

- A) Mn
- B) Ca
- C) K
- D) Cr
- E) V

Answer: D

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle);

Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

47) What element is represented by the electron configuration

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^{10}$?

- A) Ag
- B) Rb
- C) Cd
- D) Sr
- E) Cu

Answer: A

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

48) What is the electron configuration for tungsten?

- A) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^6$
- B) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^1 4f^{14} 5d^5$
- C) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^4$
- D) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^7$
- E) $1s^2 2s^2 2p^6 3s^2 3p^5 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^7$

Answer: C

Difficulty: 2 Medium

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

49) What is the electron configuration for silicon?

- A) $1s^2 2s^2 2p^6 3s^1 3p^3$
- B) $1s^2 2s^2 2p^6 3s^2 3p^2$
- C) $1s^2 2s^2 2p^6 3s^4$
- D) $1s^2 2s^2 2p^6 3p^4$
- E) $1s^2 2s^2 2p^6 3s^2 3p^3$

Answer: B

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

50) What is the electron configuration for bromine?

- A) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 4d^{10} 4p^6$
- B) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$
- C) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10} 4p^6$
- D) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$
- E) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$

Answer: E

Difficulty: 1 Easy

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle);

Energy-Level Splitting (Zeff and Shielding)

Accessibility: Keyboard Navigation

Gradable: automatic

51) Which of the following elements has the largest number of unpaired electrons in the ground state?

- A) K
- B) V
- C) S
- D) Si
- E) Cl

Answer: B

Difficulty: 3 Hard

Topic: Quantum Theory and Atomic Structure; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Quantum Numbers; Assigning Electrons to Atomic Orbitals (Aufbau Principle); Hund's Rule; Energy-Level Splitting (Zeff and Shielding); Diamagnetism and Paramagnetism

Accessibility: Keyboard Navigation

Gradable: automatic

52) The general electron configuration for atoms of all elements in Group 5A is

- A) $ns^{2np}6$.
- B) $ns^{2np}5$.
- C) $ns^{2np}4$.
- D) $ns^{2np}3$.
- E) $ns^{2np}1$.

Answer: D

Difficulty: 1 Easy

Topic: Components of Matter; Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Periodic Classification of the Elements; Elements and the Periodic Table

Accessibility: Keyboard Navigation

Gradable: automatic

53) Which of these choices is the general electron configuration for the outermost electrons of elements in the alkaline earth group?

- A) ns^1
- B) ns^2
- C) ns^2np^4
- D) ns^2np
- E) $ns^2np^6 (n - 1)d^6$

Answer: B

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

54) The general electron configuration for atoms of the halogen group is

- A) ns^2np^6 .
- B) ns^2np^5 .
- C) $ns^2np^6 (n - 1)d^7$.
- D) ns^1 .
- E) ns^2np^7 .

Answer: B

Difficulty: 1 Easy

Topic: Components of Matter; Electron Configuration; Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements; Elements and the Periodic Table

Accessibility: Keyboard Navigation

Gradable: automatic

55) The general electron configuration for noble gas atoms is

- A) ns^2np^6 .
- B) ns^2np^5 .
- C) ns^2np^4 .
- D) ns^2np^3 .
- E) ns^2 .

Answer: A

Difficulty: 1 Easy

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

56) Each of the noble gases has a completely filled p subshell except for which one?

- A) Xenon
- B) Neon
- C) Radon
- D) Argon
- E) Helium

Answer: E

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

57) An element with the general electron configuration for its outermost electrons of ns^2np^1 would be in which element group?

- A) 2A
- B) 3A
- C) 4A
- D) 5A
- E) 8A

Answer: B

Difficulty: 1 Easy

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

58) In what group of the periodic table is the element with the electron configuration $[\text{Ar}]4s^23d^{10}4p^3$?

- A) 1A
- B) 2A
- C) 3A
- D) 4A
- E) 5A

Answer: E

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

59) Consider the element with the electron configuration $[\text{Kr}]5s^24d^7$. This element is

- A) a halogen.
- B) a transition metal.
- C) a nonmetal.
- D) an actinide element.
- E) a noble gas.

Answer: B

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity; Transition Metals and Coordination Compounds

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements; Properties of Transition Metals

Accessibility: Keyboard Navigation

Gradable: automatic

60) Consider the element with the electron configuration $[\text{Kr}]5s^24d^{10}5p^5$. This element is

- A) a halogen.
- B) a transition metal.
- C) an alkali metal.
- D) an actinide element.
- E) a noble gas.

Answer: A

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

61) Consider the element with the electron configuration $[\text{Xe}]6s^24f^7$. This element is

- A) a halogen.
- B) a lanthanide element.
- C) a nonmetal.
- D) an actinide element.
- E) a noble gas.

Answer: B

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

62) How many *valence electrons* does a carbon atom have?

- A) 1
- B) 2
- C) 3
- D) 4
- E) 6

Answer: D

Difficulty: 1 Easy

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

63) How many *valence electrons* does a tin (Sn) atom have?

- A) 2
- B) 4
- C) 14
- D) 36
- E) 50

Answer: B

Difficulty: 1 Easy

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

64) An element with the electron configuration [noble gas] $ns^2(n-1)d^8$ has _____ valence electrons.

- A) 2
- B) 6
- C) 8
- D) 10
- E) None of the answers is correct.

Answer: D

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

65) An element with the electron configuration [noble gas] $ns^2(n-1)d^{10}np^3$ has _____ valence electrons.

- A) 2
- B) 3
- C) 5
- D) 10
- E) 15

Answer: C

Difficulty: 2 Medium

Topic: Electron Configuration; Chemical Periodicity

Bloom's: 4. Analyze

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Periodic Classification of the Elements

Accessibility: Keyboard Navigation

Gradable: automatic

66) How does atomic radius change as you move across the periodic table?

- A) Atomic radius decreases moving from left to right across a period and increases from top to bottom.
- B) Atomic radius increases moving left to right across a period and decreases from top to bottom.
- C) Smaller nuclear charge lowers energy; more electrons in an orbital lowers energy.
- D) Atomic radius increases diagonally across the periodic table.
- E) None of the answers is correct.

Answer: A

Difficulty: 1 Easy

Topic: Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

67) Which of these atoms has the smallest radius?

- A) Al
- B) P
- C) As
- D) Te
- E) Na

Answer: B

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

68) Which of these atoms has the largest radius?

- A) B
- B) Ga
- C) Br
- D) Si
- E) Cl

Answer: B

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

69) Which of the elements listed below has the greatest atomic radius?

- A) B
- B) Al
- C) S
- D) P
- E) Si

Answer: B

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

70) Which one of these ions has the smallest radius?

- A) Cl^-
- B) K^+
- C) S^{2-}
- D) Na^+
- E) O^{2-}

Answer: D

Difficulty: 3 Hard

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

71) Arrange P, S, and O in order of increasing atomic radius.

- A) $\text{S} < \text{O} < \text{P}$
- B) $\text{P} < \text{S} < \text{O}$
- C) $\text{O} < \text{S} < \text{P}$
- D) $\text{O} < \text{P} < \text{S}$
- E) The answer cannot be determined from the data given.

Answer: C

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

72) Arrange these ions in order of increasing ionic radius: K^+ , P^{3-} , S^{2-} , Cl^- .

- A) $\text{K}^+ < \text{Cl}^- < \text{S}^{2-} < \text{P}^{3-}$
- B) $\text{K}^+ < \text{P}^{3-} < \text{S}^{2-} < \text{Cl}^-$
- C) $\text{P}^{3-} < \text{S}^{2-} < \text{Cl}^- < \text{K}^+$
- D) $\text{Cl}^- < \text{S}^{2-} < \text{P}^{3-} < \text{K}^+$
- E) $\text{Cl}^- < \text{S}^{2-} < \text{K}^+ < \text{P}^{3-}$

Answer: A

Difficulty: 3 Hard

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

73) Which of the following elements has the largest atomic size?

- A) S
- B) Ca
- C) Ba
- D) Po
- E) Rn

Answer: C

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

74) Which of the following elements has the smallest atomic size?

- A) Na
- B) Ar
- C) K
- D) Ca
- E) Kr

Answer: B

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

75) Select the element that will lose an electron most easily, based on the periodic trend.

- A) Li
- B) Na
- C) K
- D) He

Answer: C

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

76) Select the element that will lose an electron most easily, based on the periodic trend.

- A) Na
- B) Mg
- C) Ar
- D) P

Answer: A

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

77) Select the element that will gain an electron most easily, based on the periodic trend.

- A) Ca
- B) Mg
- C) O
- D) P

Answer: C

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

78) Select the element that will gain an electron most easily, based on the periodic trend.

- A) Rb
- B) Al
- C) S
- D) Na

Answer: C

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 3. Apply

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

79) Which of these elements has the greatest metallic character?

- A) Br
- B) F
- C) Ge
- D) Mn
- E) Sc

Answer: E

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

80) Which of these elements has the greatest metallic character?

- A) Br
- B) Se
- C) Ni
- D) As
- E) Si

Answer: C

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

81) Select the element with the greatest metallic character.

- A) Li
- B) Ca
- C) Al
- D) Pb
- E) Cs

Answer: E

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

82) Select the element with the least metallic character.

- A) Sn
- B) Sr
- C) Tl
- D) Ge
- E) Ga

Answer: D

Difficulty: 2 Medium

Topic: Chemical Periodicity

Bloom's: 5. Evaluate

Subtopic: Trends in Physical Properties (Atomic Radius, Ionization Energy, Ionic Radius, Electron Affinity)

Accessibility: Keyboard Navigation

Gradable: automatic

83) Using the periodic table, predict the charge on the common ion of calcium.

- A) +1
- B) +2
- C) -1
- D) -2

Answer: B

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

84) Using the periodic table, predict the charge on the common ion of selenium.

- A) +1
- B) +2
- C) -1
- D) -2

Answer: D

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

85) Using the periodic table, predict the charge on the common ion of rubidium.

- A) +1
- B) +2
- C) -1
- D) -2

Answer: A

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

86) Using the periodic table, predict the charge on the common ion of bromine.

- A) +1
- B) +2
- C) -1
- D) -2

Answer: C

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

87) The Lewis dot symbol consists of the symbol for the element surrounded by dot(s). What does the symbol represent?

- A) electron configuration
- B) valence electrons
- C) atomic number
- D) atomic mass
- E) nucleus and core electrons

Answer: E

Difficulty: 2 Medium

Topic: Chemical Bonding; Molecular Shape

Bloom's: 2. Understand

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

88) The Lewis dot symbol consists of the symbol for the element surrounded by dot(s). What does the dot or dots represent?

- A) electron configuration
- B) valence electrons
- C) atomic number
- D) atomic mass
- E) core electrons

Answer: B

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 2. Understand

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

89) How many dots does the Lewis dot symbol for argon have around it?

- A) 1
- B) 2
- C) 4
- D) 6
- E) 8

Answer: E

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 3. Apply

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

90) How many dots does the Lewis dot symbol for sodium have around it?

- A) 1
- B) 2
- C) 0
- D) 3
- E) 7

Answer: A

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 3. Apply

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

91) How many dots does the Lewis dot symbol for magnesium have around it?

- A) 1
- B) 2
- C) 0
- D) 3
- E) 7

Answer: B

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 3. Apply

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

92) How many dots does the Lewis dot symbol for chlorine have around it?

- A) 1
- B) 2
- C) 5
- D) 7
- E) 17

Answer: D

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 3. Apply

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

93) How many dots does the Lewis dot symbol for carbon have around it?

- A) 4
- B) 2
- C) 6
- D) 3
- E) 7

Answer: A

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 3. Apply

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

94) How many dots does the Lewis dot symbol for oxygen have around it?

- A) 4
- B) 2
- C) 6
- D) 3
- E) 7

Answer: C

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 3. Apply

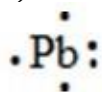
Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Accessibility: Keyboard Navigation

Gradable: automatic

95) The Lewis dot symbol for the a lead atom is

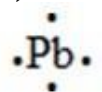
A)



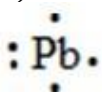
B)



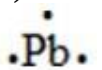
C)



D)



E)



Answer: C

Difficulty: 1 Easy

Topic: Chemical Bonding; Molecular Shape

Bloom's: 4. Analyze

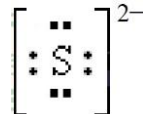
Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Gradable: automatic

96) The Lewis dot symbol for the S^{2-} ion is

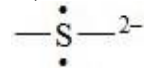


B)

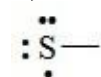


C) S^{2-}

D)



E)



Answer: B

Difficulty: 2 Medium

Topic: Chemical Bonding; Molecular Shape

Bloom's: 4. Analyze

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

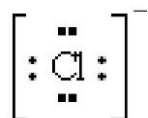
Gradable: automatic

97) The Lewis dot symbol for the chloride ion is

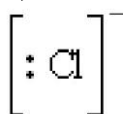
A)



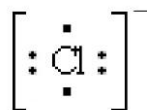
B)



C)



D)



E) Cl^{-}

Answer: B

Difficulty: 2 Medium

Topic: Chemical Bonding; Molecular Shape

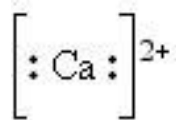
Bloom's: 4. Analyze

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Gradable: automatic

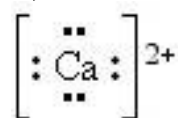
98) The Lewis dot symbol for the calcium ion is

A)



B) —Ca—

C)



D) Ca^{2+}

E) Ca

Answer: D

Difficulty: 2 Medium

Topic: Chemical Bonding; Molecular Shape

Bloom's: 4. Analyze

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Gradable: automatic

99) Select the element whose Lewis symbol is correct.

A) $\cdot \text{Fr} \cdot$

B) $\cdot \text{Ra} \cdot$

C) $\cdot \text{Pb} \cdot$

D) $: \text{Te} :$

E) $\cdot \text{He} \cdot$

Answer: B

Difficulty: 2 Medium

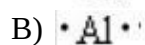
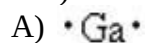
Topic: Chemical Bonding; Molecular Shape

Bloom's: 4. Analyze

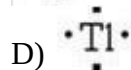
Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Gradable: automatic

100) Select the element whose Lewis symbol is correct.



C)



Answer: C

Difficulty: 2 Medium

Topic: Chemical Bonding; Molecular Shape

Bloom's: 4. Analyze

Subtopic: Lewis Dot Symbols; Writing Lewis Dot Structures

Gradable: automatic

101) A magnesium ion, Mg^{2+} , has

A) 12 protons and 13 electrons.

B) 24 protons and 26 electrons.

C) 12 protons and 10 electrons.

D) 24 protons and 22 electrons.

E) 12 protons and 14 electrons.

Answer: C

Difficulty: 2 Medium

Topic: Components of Matter; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions; Structure of the Atom

Accessibility: Keyboard Navigation

Gradable: automatic

- 102) An aluminum ion, Al^{3+} , has
- A) 13 protons and 13 electrons.
 - B) 27 protons and 24 electrons.
 - C) 16 protons and 13 electrons.
 - D) 13 protons and 10 electrons.
 - E) 10 protons and 13 electrons.

Answer: D

Difficulty: 2 Medium

Topic: Components of Matter; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions; Structure of the Atom

Accessibility: Keyboard Navigation

Gradable: automatic

- 103) An oxide ion, O^{2-} , has
- A) 8 protons and 10 electrons.
 - B) 10 protons and 8 electrons.
 - C) 8 protons and 9 electrons.
 - D) 8 protons and 7 electrons.
 - E) 10 protons and 7 electrons.

Answer: A

Difficulty: 2 Medium

Topic: Components of Matter; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions; Structure of the Atom

Accessibility: Keyboard Navigation

Gradable: automatic

- 104) A sulfide ion, S^{2-} , has
- A) 16 protons and 16 electrons.
 - B) 32 protons and 16 electrons.
 - C) 16 protons and 14 electrons.
 - D) 16 protons and 18 electrons.
 - E) 32 protons and 18 electrons.

Answer: D

Difficulty: 2 Medium

Topic: Components of Matter; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions; Structure of the Atom

Accessibility: Keyboard Navigation

Gradable: automatic

105) How many protons and electrons are present in one Br^- ion?

- A) 35 protons, 35 electrons
- B) 80 protons, 81 electrons
- C) 35 protons, 34 electrons
- D) 35 protons, 36 electrons
- E) 80 protons, 34 electrons

Answer: D

Difficulty: 2 Medium

Topic: Components of Matter; Electron Configuration

Bloom's: 4. Analyze

Subtopic: Electron Configurations of Cations and Anions; Structure of the Atom

Accessibility: Keyboard Navigation

Gradable: automatic

106) An isoelectronic series is

- A) a series that has two or more species that have identical nuclear charges, but have different electron configurations.
- B) a series that has the same ionization potentials.
- C) a series that can have only up to three species and have similar electron configuration and similar nuclear charges.
- D) a series that has two or more species that have identical electron configurations, but different nuclear charges.
- E) a series that has the same nuclear charge.

Answer: D

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 2. Understand

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle)

Accessibility: Keyboard Navigation

Gradable: automatic

107) Which of these species make an *isoelectronic pair*: Cl^- , O^{2-} , F, Ca^{2+} , Fe^{3+} ?

A) Ca^{2+} and Fe^{3+}

B) O^{2-} and F

C) F and Cl^-

D) Cl^- and Ca^{2+}

E) None of the species are part of an isoelectronic series.

Answer: D

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 5. Evaluate

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

108) Which of these pairs consists of *isoelectronic* species?

A) Mn^{2+} and Ar

B) Zn^{2+} and Cu^{2+}

C) Na^+ and K^+

D) Cl^- and S

E) K^+ and Cl^-

Answer: E

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 5. Evaluate

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

109) Which ion is *isoelectronic* with Ar?

- A) Fe^{2+}
- B) F^-
- C) Br^-
- D) Ga^{3+}
- E) Ca^{2+}

Answer: E

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 5. Evaluate

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

110) Which one of these ions is *not* isoelectronic with Kr?

- A) As^{3+}
- B) Se^{2-}
- C) Rb^+
- D) Sr^{2+}
- E) Br^-

Answer: A

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 5. Evaluate

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

111) Which of these choices is the electron configuration for the aluminum ion?

- A) $1s^2 2s^2 2p^6 3s^2$
- B) $1s^2 2s^2 2p^6 3s^2 3p^2$
- C) $1s^2 2s^2 2p^6 3s^2 3p^1$
- D) $1s^2 2s^2 2p^6$
- E) $1s^2 2s^2 2p^6 3s^2 3p^4$

Answer: D

Difficulty: 3 Hard

Topic: Electron Configuration

Bloom's: 5. Evaluate

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic

112) Which of these choices is the electron configuration for the chloride ion?

- A) $[\text{Ne}]3s^2 3p^4$
- B) $[\text{Ne}]3s^2 3p^7$
- C) $[\text{Ar}]$
- D) $[\text{Ar}]4s^1$
- E) $[\text{Ne}]3s^2 3p^5$

Answer: C

Difficulty: 2 Medium

Topic: Electron Configuration

Bloom's: 5. Evaluate

Subtopic: Assigning Electrons to Atomic Orbitals (Aufbau Principle); Electron Configurations of Cations and Anions

Accessibility: Keyboard Navigation

Gradable: automatic